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$\text{Zn(s)} + \text{Cu}^{2+} \rightarrow \text{Zn}^{2+} +$

Cu(s) An electrolytic

cell based on the

reaction repre- sented

above was constructed

from zinc and copper

half-cells. The observed

voltage was found to be

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1.00 volt instead of the standard cell potential, E° , of 1.10 volts. Which of the following could correctly account for this observation?

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$\text{MgF}_2(\text{s}) \rightleftharpoons$

$\text{Mg}^{2+}(\text{aq}) + 2 \text{F}^{-}(\text{aq})$ In a
saturated solution of

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MgF₂ at 18 °C, the concentration of Mg²⁺ is 1.21 × 10⁻³ molar. The equilibrium is represented by the equation above. <https://www.scribd.com/document/127062936/AP-Chemistry-1994-Free-Response> read more

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of water decreases while

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10% sugar solution and

when equilibrium is

reached the water will

evaporate and the

solution will increase in

volume.

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principle of atomic
structure and/or
chemical bonding to
answer of each of the
following. (a) The
radius of the Ca atom is
0.197 nanometer; the
radius of the Ca^{2+} ion
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a) Potassium has a lower first-ionization energy than lithium.

b) The ionic radius of N^{3-} is larger than that of O^{2-} .

c) A calcium atom is larger than a zinc atom.

d)

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