

## Chemistry Chemical Equations Study Guide

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Construct and revise an explanation for the outcome of a simple chemical reaction based on the outermost electron states of atoms, trends in the periodic table, and knowledge of the patterns of chemical properties. Use mathematical representations to support the claim that atoms, and therefore mass, are conserved during a chemical reaction.

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order to have the necessary foundation to build upon as you continue in the Chemistry concentration in ABE. It is very important to note that this course is a pre-requisite to all the other ABE Chemistry courses.

## Chemical Reactions Study Guide

Sample Study Sheet 4.1: Balancing Chemical Equations describes the important skill of balancing equations so that they reflect the fact that the number of atoms of each element for the products of chemical changes always equals the number of atoms of each element in the initial reactants.

## Chapter 4 An Introduction to Chemical Reactions

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## Chemistry Chemical Equations Study Guide

Balancing a chemical equation can take some skill, but we'll give you plenty of practice later. To continue our sodium chloride synthesis example, the correct and balanced chemical equation is:  $2 \text{Na} (\text{s}) + \text{Cl}_2 (\text{g}) \rightarrow 2 \text{NaCl} (\text{s})$  The large numbers that are placed in front of chemical formulas are called coefficients.

## Chemical Equations Help | Chemical Reactions Study Guide ...

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change Identify the reactants and products in a chemical reaction Represent chemical reactions with equations Know what is represented by the symbols: s, l, g, and aq Categorize chemical

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For most chemical reactions in this book, solids are labeled (s), liquids (l), and gases (g). The numerical coefficients in front of the chemical formulas express the moles of each compound or element. The preceding reaction can be interpreted in terms of moles or masses. (See Table 1.)

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Answer and Explanation: Given- The given chemical reaction equation is  $\text{NH}_3 \rightarrow \text{N}_2 + \text{H}_2$  . The step by step balancing of the given chemical reaction equation is shown below,  $\text{NH}_3$  ...

[Balance the following chemical equation:  \$\text{NH}\_3\$  ... - study.com](#)

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